

## STUFF I SHOULD KNOW FOR THE AP TEST BUT DO NOT KNOW YET

### IONS LIST

acetate	$\text{C}_2\text{H}_3\text{O}_2^-$	ferric	$\text{Fe}^{3+}$	oxalate	$\text{C}_2\text{O}_4^{2-}$
aluminum	$\text{Al}^{3+}$	ferrous	$\text{Fe}^{2+}$	oxide	$\text{O}^{2-}$
ammonium	$\text{NH}_4^+$	fluoride	$\text{F}^-$	perbromate	$\text{BrO}_4^-$
barium	$\text{Ba}^{2+}$	hydrogen	$\text{H}^+$	perchlorate	$\text{ClO}_4^-$
bicarbonate	$\text{HCO}_3^-$	hydrionium	$\text{H}_3\text{O}^+$	periodate	$\text{IO}_4^-$
bisulfate	$\text{HSO}_4^-$	hydroxide	$\text{OH}^-$	permanganate	$\text{MnO}_4^-$ (purple)
bisulfide	$\text{HS}^-$	hypobromite	$\text{BrO}^-$	peroxide	$\text{O}_2^{2-}$
bisulfite	$\text{HSO}_3^-$	hypochlorite	$\text{ClO}^-$	phosphate	$\text{PO}_4^{3-}$
bromate	$\text{BrO}_3^-$	hypoiodite	$\text{IO}^-$	phosphide	$\text{P}^{3-}$
bromide	$\text{Br}^-$	iodate	$\text{IO}_3^-$	phosphite	$\text{PO}_3^{3-}$
bromite	$\text{BrO}_2^-$	iodide	$\Gamma$	potassium	$\text{K}^+$
calcium	$\text{Ca}^{2+}$	iodite	$\text{IO}_3^-$	silver	$\text{Ag}^+$
carbonate	$\text{CO}_3^{2-}$	lead	$\text{Pb}^{2+}$	sodium	$\text{Na}^+$
chlorate	$\text{ClO}_3^-$	lithium	$\text{Li}^+$	stannic	$\text{Sn}^{4+}$
chloride	$\text{Cl}^-$	magnesium	$\text{Mg}^{2+}$	stannous	$\text{Sn}^{2+}$
chlorite	$\text{ClO}_2^-$	manganese	$\text{Mn}^{2+}$	strontium	$\text{Sr}^{2+}$
chromate	$\text{CrO}_4^{2-}$ (yellow)	mercuric	$\text{Hg}^{2+}$	sulfate	$\text{SO}_4^{2-}$
chromium	$\text{Cr}^{3+}$	mercurous	$\text{Hg}_2^{2+}$	sulfide	$\text{S}^{2-}$
cupric	$\text{Cu}^{2+}$ (blue)	nickel	$\text{Ni}^{2+}$ (green)	sulfite	$\text{SO}_3^{2-}$
cuprous	$\text{Cu}^+$ (blue)	nitrate	$\text{NO}_3^-$	thiocyanate	$\text{SCN}^-$
cyanide	$\text{CN}^-$	nitride	$\text{N}^{3-}$	thiosulfate	$\text{S}_2\text{O}_3^{2-}$
dichromate	$\text{Cr}_2\text{O}_7^{2-}$ (orange)	nitrite	$\text{NO}_2^-$	zinc	$\text{Zn}^{2+}$

### SOLUBILITY RULES

Always soluble:

alkali metal ions ( $\text{Li}^+$ ,  $\text{Na}^+$ ,  $\text{K}^+$ ,  $\text{Rb}^+$ ,  $\text{Cs}^+$ ),  $\text{NH}_4^+$ ,  $\text{NO}_3^-$ ,  $\text{ClO}_3^-$ ,  $\text{ClO}_4^-$ ,  $\text{C}_2\text{H}_3\text{O}_2^-$

Generally soluble: (mnemonics)

$\text{CT}^-$ ,  $\text{Br}^-$ ,  $\Gamma^-$  Soluble except  $\text{Ag}^+$ ,  $\text{Pb}^{2+}$ ,  $\text{Hg}_2^{2+}$  (AP/H)  
 $\text{F}^-$  Soluble except  $\text{Ca}^{2+}$ ,  $\text{Sr}^{2+}$ ,  $\text{Ba}^{2+}$ ,  $\text{Pb}^{2+}$ ,  $\text{Mg}^{2+}$  (CBS-PM)

$\text{SO}_4^{2-}$  Soluble except  $\text{Ca}^{2+}$ ,  $\text{Sr}^{2+}$ ,  $\text{Ba}^{2+}$ ,  $\text{Pb}^{2+}$  (CBS/PBS)

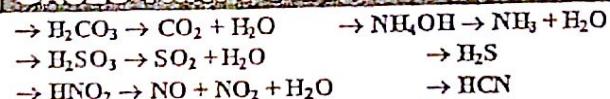
Generally insoluble:

$\text{O}^{2-}$ ,  $\text{OH}^-$  Insoluble except and alkali metals, and  $\text{NH}_4^+$   
 $\text{Ca}^{2+}$ ,  $\text{Sr}^{2+}$ ,  $\text{Ba}^{2+}$  (CBS) somewhat soluble

$\text{CO}_3^{2-}$ ,  $\text{PO}_4^{3-}$ ,  $\text{S}^{2-}$ ,  $\text{SO}_3^{2-}$ ,  $\text{C}_2\text{O}_4^{2-}$ ,  $\text{CrO}_4^{2-}$

Insoluble except alkali metals and  $\text{NH}_4^+$

### GASES THAT FORM



### WEAK ELECTROLYTES

Weak Acids (esp.  $\text{HC}_2\text{H}_3\text{O}_2$  and HF)

(Memorize the 8 strong acids... all others are weak)

HCl	hydrochloric acid	$\text{HNO}_3$	nitric acid
HBr	hydrobromic acid	$\text{HIO}_4$	periodic acid
HI	hydroiodic acid	$\text{H}_2\text{SO}_4$	sulfuric acid
$\text{HClO}_4$	perchloric acid	$\text{HClO}_3$	chloric acid

Ammonium Hydroxide ( $\text{NH}_4\text{OH} \approx \text{NH}_3(\text{aq})$ ) Water ( $\text{H}_2\text{O}$ )

### DRIVING FORCES — Double Replacement

- Insoluble Solid (Precipitate)
- Weak Electrolyte ( $\text{H}_2\text{O}$  or Weak Acid)
- Gas Formation

### STRONG OXIDIZERS (Oxidizing Agents)

- ✗  $\text{MnO}_4^-$  in acid solution  $\rightarrow \text{Mn}^{2+} + \text{H}_2\text{O}$
- ✗  $\text{MnO}_2$  in acid solution  $\rightarrow \text{Mn}^{2+} + \text{H}_2\text{O}$
- ✗  $\text{MnO}_4^-$  in neutral or basic sol'n  $\rightarrow \text{MnO}_2$
- ✗  $\text{Cr}_2\text{O}_7^{2-}$  in acid solution  $\rightarrow \text{Cr}^{3+} + \text{H}_2\text{O}$
- ✗  $\text{Cr}_2\text{O}_7^{2-}$  with a base  $\rightarrow \text{CrO}_4^{2-} + \text{H}_2\text{O}$
- ✗  $\text{CrO}_4^{2-}$  in basic solution  $\rightarrow \text{CrO}_2^- + \text{H}_2\text{O}$
- ✗  $\text{HNO}_3$ , concentrated  $\rightarrow \text{NO}_2 + \text{H}_2\text{O}$
- ✗  $\text{HNO}_3$ , dilute (e.g. 6 M)  $\rightarrow \text{NO} + \text{H}_2\text{O}$
- ✗  $\text{H}_2\text{SO}_4$ , hot, concentrated  $\rightarrow \text{SO}_2 + \text{H}_2\text{O}$
- ✗ Free halogens (e.g.  $\text{Cl}_2$ )  $\rightarrow$  halide ions ( $\text{Cl}^-$ )
- ✗  $\text{H}_2\text{O}_2$  in acid solution  $\rightarrow \text{H}_2\text{O}$
- ✗ Note:  $\text{H}_2\text{O}_2$  decomposes  $\rightarrow \text{H}_2\text{O} + \text{O}_2$
- ✗  $\text{Na}_2\text{O}_2$   $\rightarrow \text{NaOH}$
- ✗  $\text{HClO}_4$   $\rightarrow \text{Cl}^- + \text{H}_2\text{O}$

#### Other Oxidizers

- ✗ Metal-“ic” ions (e.g.  $\text{Sn}^{4+}$ ,  $\text{Fe}^{3+}$ )  $\rightarrow$  “ous” ions ( $\text{Sn}^{2+}$ ,  $\text{Fe}^{2+}$ )
- ✗  $\text{H}_2\text{O}$   $\rightarrow \text{H}_2 + \text{OH}^-$

### STRONG REDUCERS (Reducing Agents)

- ✗ Halide ions (e.g.  $\text{Cl}^-$ )  $\rightarrow$  Free halogen ( $\text{Cl}_2$ )
- ✗ Free metals  $\rightarrow$  metal ions
- ✗ “ites”  $\text{SO}_3^{2-}$  or  $\text{SO}_2$ ,  $\text{NO}_2^-$   $\rightarrow$  “ates”  $\text{SO}_4^{2-}$ ,  $\text{NO}_3^-$
- ✗ Free halogens, dil. basic sol'n  $\rightarrow$  hypohalite ions ( $\text{ClO}^-$ )
- ✗ Free halogens, conc. basic sol'n  $\rightarrow$  halate ions ( $\text{ClO}_3^-$ )
- ✗  $\text{S}_2\text{O}_3^{2-}$   $\rightarrow \text{S}_4\text{O}_6^{2-}$

#### Other Reducers

- ✗ Metal-“ous” ions (e.g.  $\text{Sn}^{2+}$ )  $\rightarrow$  “-ic” ions ( $\text{Sn}^{4+}$ )
- ✗  $\text{H}_2\text{O}$   $\rightarrow \text{O}_2 + \text{H}^+$